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Chapter 10 Gases

Chapter 11 Liquids and Intermolecular Forces

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unit area on a surface. Newton (N) The force

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that will increase the speed of a one-kilogram mass by one meter per second each second that the force is applied. Barometer. A device used to measure atmospheric pressure.

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instrument that measures atmospheric pressure.

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Pressure (P) is defined as the force per unit of area on a surface. $\text{pressure} = \text{force} \div \text{area}$
 $P = F/A$ Atmospheric pressure (atm) is the pressure exerted on an object due to the weight of the column of the air above it in the atmosphere.

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; Gases Pressure; 2 Pressure Force. Pressure (P) is defined as the force per unit of area on a surface. $P = F/A$; Atmospheric pressure (atm) is the pressure exerted on an object due to the weight of the column of the air above it in the atmosphere.

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An ideal gas adheres exactly to the kinetic theory of gases. 11.3: Pressure - The Result of Constant Molecular Collisions Pressure is a force exerted over an area. Pressure has several common units that can be converted. 11.4: The Combined Gas Law- Pressure, Volume, and Temperature

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Chapter 11 185 Reread the Study Sheets in this chapter and decide whether you will use them or some variation on them to complete

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the tasks they describe. Sample Study Sheet
11.1: Using the Ideal Gas Equation Sample
Study Sheet 11.2: Using the Combined Gas Law
Equation Sample Study Sheet 11.3: Equation
Stoichiometry Problems

Chapter 11 - Gases - An Introduction to Chemistry

Chapter 11 - Gases. Chapter 11 of Modern
Chemistry Textbook. STUDY. PLAY. pressure.
the force applied to a unit area of surface.
newton. a unit of force equal to the force
that imparts an acceleration of 1 m/sec/sec
to a mass of 1 kilogram. barometer. An

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instrument used to measure air pressure.

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About This Chapter The Gases chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with gases. Each of these simple and fun video...

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CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Pressure surf f a o c r e

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ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. Rank the following pressures in increasing order. (a) 50 kPa (c) 76 torr (b) 2 atm (d) 100 N/m² 3.

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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as $P_{\text{CO}_2} = 0.285 \text{ torr}$ and $P_{\text{N}_2} = 593.525 \text{ torr}$.

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What is the partial pressure of oxygen?

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Chapter 11 Review Gases Section 2 Answers

Modern Chemistry ideal gases. To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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