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Answers

Answer: The average kinetic energy of the molecule is $KE = \left(\frac{1}{2} m$

$\bar{v}^2\right) \dots\dots\dots(3)$ The eq (2) can

be modified as. The average Kinetic energy of a molecule remains constant when the temperature is constant.

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Plus One Physics Notes Chapter 13

Kinetic Theory - HSSLive

Kinetic Theory The kinetic theory is an explanation of how particles in matter behave. To explain the behavior of particles, it is necessary to make some basic assumptions. The three assumptions of the kinetic theory are as follows: 1. All

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Answers
matter is composed of small particles (atoms, molecules, and ions). 2. These particles are in constant, random motion. 3.

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Answers on the idea that the gas consists of rapidly moving atoms or molecules.

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The host describes the different properties

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Answers of gases and the components of Kinetic Molecular Theory. Students are asked to make predictions about what will happen during demonstrations with gases using a balloon and marshmallow, applying the concepts of ideal gas law. They are asked to create a model demonstrating what happened at the molecular level.

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