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The stoichiometry of a
reaction describes the

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relative amounts of reactants and products in a balanced chemical equation. A stoichiometric quantity of a reactant is the amount necessary to react completely with the other reactant (s). If a reactant

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remains unconsumed after
complete reaction has
occurred, it is in excess.

~~3: Stoichiometry of Formulas
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Stoichiometry. study of quantitative relationships between the amounts of reactants used and the amounts of products formed in a chemical reaction.

Conduct two separate mass-to-mass (or mole-to-mole)

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stoichiometric calculations;
comparing the mass of each
reactant to a chosen product

...

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~~Chapter 11 Stoichiometry ...~~

Stoichiometry involves the
concept of the mole.

Particulate matter and
chemical change Writing
formulae and chemical
equations is a fundamental
skill of a chemist (it is

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(also a universal language).

You are probably familiar with balancing chemical equations and many of the key terms in this part of the course.

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In ideal Stoichiometry - All reactants are converted into products. Common Methods for solving all Stoichiometry Problems: Mass-Mass

Problems: 1. Start with a

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known mass of reactant or product and find an unknown mass of another reactant or product. 2. All other stoichiometry problems are derivations (shortened versions) of this larger solution:

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established from the law of conservation of mass wherein it correlates the reactants and the products quantitatively. Chemical reactions have a reactant side and product side.

The... Why is stoichiometry

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