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Solubility and Balancing Reactions

Predicting Products More Examples

and Practice: How to Predict and

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Valence Electrons and the Periodic

~~Table~~ Types of Chemical Reactions

Writing Formulas from Names (1 of 2)

Molecular, Ionic, and Net Ionic

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Experiment 1 Chemical Reactions and
Net Ionic Equations I. Objective: To
predict the products of some
displacement reactions and write net

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ionic equations. II. Chemical Principles
: A. Reaction Types. Chemical
reactions can be classified into several
different categories. The common
classifications are as follows:

Experiment 1 Chemical Reactions and Net Ionic Equations

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Write ionic equations for the following:

1. The neutralisation reaction between hydrochloric acid and barium hydroxide
2. The neutralisation reaction between sulfuric acid and ammonium hydroxide
3. The precipitation

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reaction between iron(II) sulfate and sodium hydroxide 4. The precipitation reaction between lead(II) nitrate and ammonium sulfate 5.

9-1 GCSE Chemistry Ionic Equations Questions

Since the reaction of HCl and NaOH

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occurs in a 1:1 ratio, at the

equivalence point of the titration, the moles of NaOH added must be equal to the moles of HCl in solution. The reaction equation is. $\text{HCl (aq)} + \text{NaOH (aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{Cl}^- \text{(aq)} + \text{H}_2\text{O (l)}$

(l) In every titration, we need a way to determine the point at which the

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reaction is complete.

Ions in Solution (Worksheet) -
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Net Ionic Equation Worksheet. READ
THIS: When two solutions of ionic
compounds are mixed, a solid may
form. This type of reaction is called a

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precipitation reaction, and the solid produced in the reaction is known as the precipitate. You can predict whether a precipitate will form using a list of solubility rules such as those found in the table below.

Net Ionic Equation Worksheet

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Net Ionic Equation: $2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+} (\text{aq}) \rightarrow \text{PbCl}_2 (\text{s})$ Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution. Include states of matter in your

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balanced equation. 1. Sodium chloride
and lead II nitrate Molecular Equation:
Complete Ionic Equation:

Net Ionic Equation Worksheet
Answers

Net Ionic Equation Worksheet
Answers. Write balanced molecular,

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Solubility and net ionic equations (NIE) for each of the following reactions.

Assume all reactions occur in aqueous solution.

1. $2\text{NaCl}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$

Ionic Equation: $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{NO}_3^-(\text{aq})$

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Net Ionic Equation Worksheet Answers

Write the balanced chemical reaction for reaction 1. Be sure to include the states of matter and correct formulas. If applicable (ie: if you classified it as a precipitation reaction) include the net ionic equation as well. Write the

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balanced chemical reaction for reaction 3. Be sure to include the states of matter and correct formulas.

Lab 6 Worksheet | Chemistry I
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PRACTICE PROBLEMS ON NET IONIC
EQUATIONS page 1 of 3 Show the

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complete ionic and net ionic forms of the following equations. If all species are spectator ions, please indicate that no reaction takes place. Note: you need to make sure the original equation is balanced before proceeding! A set of solubility rules are given at the end of this ...

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Chem Worksheet 8-2 Name An ionic compound is a combination of oppositely charged ions. Ionic compounds generally contain a metal bonded to a non-metal (or non-metals). When naming ionic compounds we simply name the cation (the positive ion) then the

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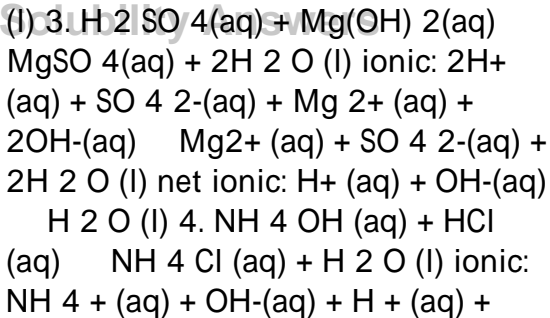
anion (the negative ion). The cations generally retain the name of the element.

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ionic: $\text{H}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) + \text{K}^+ (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{K}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) + \text{H}_2\text{O} (\text{l})$

net ionic: $\text{H}^+ (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{H}_2\text{O}$

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Cl⁻(aq) + NH₄⁺(aq) + Cl⁻(aq) + H₂O (l)
net ionic: H⁺(aq) + OH⁻(aq) ...

9-1 GCSE Chemistry Ionic Equations Answers

Writing the molecular, ionic, and net ionic equations for a chemical reaction in aqueous solution: The

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steps for writing a net ionic equation for the metathesis reaction of aqueous K_2SO_4 and $Pb(C_2H_3O_2)_2$ are outlined below. 1. Write the balanced molecular equation which gives the complete chemical formula and the phase of ...

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By mixing sodium hydroxide, NaOH (aq), with acetic acid, HC₂H₃O₂ (aq), no reaction precipitate is observed due to the formation of NaC₂H₃O₂ (aq) which is soluble according to the solubility rule. The reaction can be written as NaOH (aq) + HC₂H₃O₂ (aq) —> NaC₂H₃O₂ (aq) + H₂O (l)

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If you can ' t tell an alkali metal from
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Part 2 Practice Worksheet - write the

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complete ionic equation . Note: this is the same worksheet as in part 1. You should take your answers from the worksheet in part 1, and now rewrite them as complete ionic equations. This means you need to redo the reactants if they are soluble, and so you may want to write this on a

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separate sheet of paper.

3.5: Precipitation Reactions - Chemistry LibreTexts

This AP Chemistry class covers Topics
4.1-4.4. 4.1 Introduction to Reactions;
4.2 Net Ionic Equations; 4.3
Representations of Reactions; 4.4

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AP Chemistry: 4.1-4.4 Reactions, Net
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