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Answers

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SHORT ANSWER Answer the following questions in the space

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provided. 1. Given the following equation: $C_3H_4(g) + x \cdot O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What ...

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Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the

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following equation: $N_2(g) + 3H_2(g) \dots$

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ANSWER Answer the following questions in the space provided. 1.

Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

_____ a. What is the value of the coefficient x in this equation?

40.07 g/mol h. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol

H_2O c. What is the mole ratio of O_2 to H_2O in the above equation?

0.20 mol d. How many moles ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS

Write the answer on the line to the left Show all your work in the space provided

1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield

2 60 mol of N₂ are mixed with 120 mol of H₂ Read Online Chapter 9 Page 8/15

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