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# **Chapter 4 Review Arrangement Of Electrons In Atoms Answer Key**

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SHORT ANSWER Answer the

following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the

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Answer the following questions in the  
space provided. 1. State the Pauli

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exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states.

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Answer the following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states.

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CHAPTER 4: ARRANGEMENT OF

ELECTRONS IN ATOMS. The following pages contain the bulk (but not all) of the information for the chapter 4 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If



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CHAPTER 4 REVIEW . Arrangement  
of Electrons in Atoms . SHORT  
ANSWER Answer the following  
questions in the space provided. 1.

1. How many quantum numbers  
are used to describe the properties of

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electrons in atomic orbitals? (a) 1 (c) 3  
(b) 2 (d) 4 . 2. \_a,--\_ A spherical  
electron cloud surrounding an atomic  
nucleus would best represent

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15 thru 22. Section 3: Chapter review  
26 thru 38 . Homework Answers.  
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Electrons in Atoms states that a maximum of two electrons can occupy a single atomic orbital but only if the electrons have opposite spins Hund's rule states that single electrons with the same spin must occupy each equal-energy orbital before additional

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electrons with opposite spins can  
occupy the same orbitals

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