

Chapter 25 Nuclear Chemistry Ppt

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ppt on hydrogen for class XI th chemistry 25. The overall reaction produces 2.3×10^3 joules for every gram of carbon produced at 650°C . Reaction temperatures are in the range of 450 to 600°C Heavy water is used as a neutron moderator and coolant for nuclear reactors. Deuterium is also a potential fuel for commercial nuclear fusion 2H

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$\text{BE} + 19\text{F} + 101\text{n} \rightarrow \text{binding energy per nucleon} = \frac{\text{binding energy}}{\text{number of nucleons}}$
 $2.37 \times 10^{-11} \text{ J} / 19 \text{ nucleons} = \text{Nuclear binding energy (BE) is the energy required to break up a nucleus into its component protons and neutrons.}$
 $E = mc^2$
 $\text{BE} = 9 \times (\text{p mass}) + 10 \times (\text{n mass}) - 19\text{mass}$
 $\text{BE (amu)} = 9 \times 1.007825 + 10 \times 1.008665 \dots$

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Radiochemistry and Nuclear Chemistry

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