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Reactions Nature of Oxidation  
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Explain in your own words the

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following oxidation processes: a. Gain of oxygen b. Loss of hydrogen c. Loss of electrons 2. Explain in your own words the following reduction processes: a. Loss of oxygen b. Gain of hydrogen c. Gain of electrons 3.

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Oxidation - Reduction Rxns 1  
Section 22.1 The Meaning of  
Oxidation and Reduction 2  
OBJECTIVES: - Define the terms  
"oxidation" and "reduction" in  
terms of electron loss or gain. -  
State the characteristics of a  
RedOx Rxn, and recognize  
oxidizing and reducing agents.

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except all students skip "Using Oxidation Number Change" pp 663-664, and Regents students can skip section 23.2 "Half-Cells and Cell Potentials" pp685-691.

□ Lab 24: Oxidation-Reduction Reactions

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Reduction Reactions Section 20.1  
– Oxidation vs. Reduction

reactions are also known as  
reactions. is the of electrons or  
the gain of. is the of electrons or  
the loss of. The way to remember  
the difference in oxidation and  
reduction is.

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Chapter 20 Outline.docx - Chapter  
20 Outline Oxidation ...

! 207! Chapter12:!OxidationandR  
eduction.!! Oxidation)reduction(re  
dox)reactions. At!different!times,!  
oxidation!and!reduction!(redox)!h  
avehaddifferent,but ...

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Oxidation)reduction(redox)reactio  
ns.

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Recognize a reaction as an oxidation-reduction reaction.

Consider this chemical reaction:  $2 \text{Na (s)} + \text{Cl}_2(\text{g}) \rightarrow 2 \text{NaCl}$ . The reactants are elements, and it is assumed that they are electrically neutral; they have the same number of electrons as protons. The product, however, is ionic; it is composed of  $\text{Na}^+$  and  $\text{Cl}^-$  ions.

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Introductory Chemistry ...

In this video I'll teach you about several oxidation and reduction reactions, which include carbonyl reductions using sodium borohydride, lithium aluminum hydride, and DIBAL-H (diisobutyl aluminum ...

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Reduction Reactions: Part 1 of 2  
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oxidation reduction reactions also  
known as redox reactions. This  
video explains how to ide...

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Definition Redox reaction. It is a reduction - oxidation reaction that occurs between a reducing agent and an oxidizing agent. Ox 1 + Red 2 Red 1 + Ox 2 Where: Ox1 - is an oxidizing agent and Red2 - is a reducing agent.

Example:  $I_2 + Sn^{2+} = 2I^- + Sn^{4+}$   
oxidizing agent    reducing agent  
being oxidized    being reduced

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CHAPTER 4 TITRIMETRIC

*Page 14/15*

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ANALYSIS - redox titration pdf ...

This means that ethanol is oxidized more readily than hydrogen (so it is a better reducing agent, it releases its electrons more easily), and allows us to predict the direction of the redox reaction  $\epsilon_o = 0 - (-0.197) = +0.197$ . So electrons flow from ethanol to H + Reduction potentials and redox reactions (Chapter 3.9)