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Chapter 15 Acids and Bases. STUDY.
PLAY. Chapter 15 main idea. acids and
bases change the color of compound
indicators. ... -Lewis acid-base
reaction is the formation of one or

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more covalent bonds between an electron-pair donor and an electron pair acceptor, although the 3 acid-base reactions differ, many compounds can be categorized as acids ...

~~Chapter 15 Acids and Bases~~

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Chapter 15: Acids and Bases Chapter
15: Acids and Bases 1. The hydronium
ion and the hydroxide ion, in that
order, are: A) H_3O^+ , OH^- B) OH^- , H_3O^+
C) OH^- , H^+ D) H_3O^+ , OH^- E) H_3O^- , OH^- Ans: D Category: Easy Section:
15.1 2. Which of the following does

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not fit the definition of a Brønsted
Acid?

~~Chapter 15 Acids and Bases~~

~~Review.doc Chapter 15 Acids ...~~

Chapter 15 Acids and Bases. strong
acid. strong base. weak acid. weak
base. an acid that ionizes completely

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~~Base Answer Key~~
in solvent. a base that ionizes completely in a solvent. an acid that releases few hydrogen ions in aqueous solution. a base that releases few hydroxide ions in aqueous solution.

~~Chapter 15 Review Acids Bases 2~~

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ACIDS AND BASES 453 SECTION 15-1

OBJECTIVES List five general properties of aqueous acids and bases. Name common binary acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory,

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Base Answer Key and give two properties of each.

Define acid and base according to Arrhenius ' s theory of ionization.

Explain the differences

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Chapter 15: Acids and Bases Acids
and Bases Arrhenius Definitions:
acids - compounds that produce an

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increase in $[H^+]$ when dissolved in water bases - compounds that produce an increase in $[OH^-]$ when dissolved in water Lewis Definitions: acids - electron pair acceptors bases - electron pair donors Brønsted-Lowry Definitions: acids - H^+ donors

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Thus, H_2O and H_3O^+ , and H_2O and OH^- are conjugate acid-base pairs, but H_3O^+ and OH^- are not conjugate acid-base pair. A Brønsted-Lowry acid-base reaction involves a competition between two bases for a proton, in

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Base Answer Key
which the stronger base ends up being the most protonated at equilibrium. In the reaction: $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$,

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Base Answer Key
acid. strong base. weak acid. weak base. an acid that ionizes completely in solvent. a base that ionizes completely in a solvent. an acid that releases few hydrogen ions in aqueous solution. a base that releases few hydroxide ions in aqueous solution.

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Chapter 15 Review Acids Bases Label
them as acid, base, conjugate acid
(ca) and conjugate base (cb) Answer:
 CH_3NH_2 (aq) is the base (an amine),
 H_2O (l) acts as the acid here, CH_3NH

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~~Bases Answer Key~~
3 + (aq) is the ca, OH-(aq) is the cb;
Remember that H + is just short hand
for H 3 O + called hydronium ion. H +
doesn't really exist in water.

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acid. strong base. weak acid. weak
base. an acid that ionizes completely
in solvent. a base that ionizes
completely in a solvent. an acid that
releases few hydrogen ions in
aqueous solution. a base that releases
few hydroxide ions in aqueous
solution.

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Chapter 14 Review: Acids and Bases.
Section 14.1 Bronsted Acids and
Bases; Acids - molecules that can lose
 H^+ (proton donors) making H_3O^+
in water ... Since x is $=[H^+]$ find the

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$\text{pH} = -\log 7.0746 \times 10^{-3} = 2.15$ (we need 2 decimal places in the final answer) For more ...

~~Chapter 15 Review Acid Base
Equilibria~~

CHAPTER 14 REVIEW Acids and Bases
SECTION 1 SHORT ANSWER Answer

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the following questions in the space provided. 1. Name the following compounds as acids: sulfuric acid a. H_2SO_4 sulfurous acid b. H_2SO_3 hydrosulfuric acid c. H_2S perchloric acid d. HClO_4 hydrocyanic acid e. hydrogen cyanide 2. H

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more than additional will meet the
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answers can be taken as competently
as picked to act. Page 2/26

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