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limiting reactant. The study of quantitative relationships between the amounts of.... In a balanced equation, the ratio between the numbers of moles.... A reactant that remains after a chemical reaction stops.

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Stoichiometry comes

from the Greek words

stoikheion, which means

element, and metron,

which means to measure

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4, 4, 72.08 10. methanol  
and oxygen gas 11.  
carbon dioxide and  
water 12. 160.10 g 13.  
160.10 g 14. They are  
equal. 15. A mole ratio  
is a ratio between the  
numbers of moles

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Limiting Reactants 11.4  
Percent Yield

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In Section 11.3 , for  
example, you learned  
how to express the  
stoichiometry of the  
reaction for the  
ammonium dichromate

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14 Study  
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volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer

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questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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Stoichiometry MAIN  
Idea The amount of  
each reactant present at  
the start of a chemical

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11.1 Reaction determines how much product can form.

11.2 Stoichiometric Calculations MAIN

Idea The solution to every stoichiometric problem requires a balanced chemical

equation. 11.3 Limiting Reactants MAIN Idea A chemical reaction

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11 SOLUTIONS

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pages 368 – 372

Practice Problems pages

371 – 372 1. Interpret

the following balanced

chemical equa-tions in

terms of particles, moles,

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and mass. Show that the  
law of conservation of  
mass is

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stoichiometry? In your

textbook, read about

stoichiometry and the

balanced equation. For

each statement below,

write true or false. 2. 3.

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4.5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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