

## Acids Bases Pogil Answers

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Acids and Bases, Basic Introduction, Multiple Choice Practice Problems Chemistry

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Acids and bases are important substances in many settings. In everyday life we encounter acids such as vinegar and lemon juice and bases such as ammonia and lye in drain cleaners. Many important processes in chemistry and biology involve reactions of acids with bases ...

### Pogil Answers Acids And Bases

An Arrhenius acid is any substance that produces hydrogen ions [or hydronium ions (H<sub>3</sub>O<sup>+</sup>) a hydrogen ion attached to a water molecule] when dissolved in water. An Arrhenius base is any substance that produces hydroxide ions when dissolved in water. While the Arrhenius definitions of acids and bases is useful, it is limited.

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Many important processes in chemistry and biology involve reactions of acids with bases. Acid-base reactions can be understood in terms of the transfer of hydrogen ion, and the concentration of hydrogen ion in a reaction mixture, measured as pH, is often an important factor in how reactions occur.

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STOP Conjugate Acid  $\text{CO}_2(\text{aq}) + \text{H}_2\text{O}(\text{aq}) \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{H}^+(\text{aq})$  Model 3 Conjugate Acid-Base Pairs Base + 1-120 (l)  $\text{HC}_2\text{O}_4^-(\text{aq})$  13. All acid-base reactions have two conjugate acid-base pairs. One conjugate acid-base pair in the reaction in Model 3 is  $\text{H}_2\text{O}/\text{H}^+$ . List the other acid-base pair in the reaction. 14.

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when equal molarities of acid and conjugate base are combined? 15. When unequal molarities of acid and conjugate base are combined to make a buffer (as in beaker 4 of Model 2), does the pH change significantly from the  $\text{pK}_a$  of the weak acid? Support your answer with data from Model 2. 16. An industrial process requires a constant pH of 3.00.

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Phosphoric acid ( $\text{H}_3\text{PO}_4$ ) is a triprotic acid, but phosphorous acid ( $\text{H}_3\text{PO}_3$ ) is a diprotic acid, and hypophosphorous acid ( $\text{H}_3\text{PO}_2$ ) is a monoprotic acid. a. Draw Lewis dot-diagrams for each of these three acids.

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b.  $\text{HNO}_2$  is a weak acid, but  $\text{HNO}_3$  is a strong acid  $\text{HNO}_2$  has one terminal oxygen, but  $\text{HNO}_3$  has two, making it a stronger acid. c.  $\text{H}_3\text{PO}_4$  is a stronger acid than  $\text{H}_3\text{AsO}_4$  P is more electronegative than As, making  $\text{H}_3\text{PO}_4$  a stronger acid. d.  $\text{FCH}_2\text{CO}_2\text{H}$  is a stronger acid than  $\text{CH}_3\text{CO}_2\text{H}$  Electronegative fluorine withdraws electron density from the -OH bond of the carboxyl

### Chem 116 POGIL Worksheet - Week 10 - Solutions Weak Acid ...

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a double-stranded, helical nucleic acid molecule capable of replicating and determining the inherited traits of organisms and viruses.

Nucleotide a building block of DNA or RNA, consisting of a sugar, a nitrogenous base and a phosphate group. Double helix a pair of parallel helices intertwined about a common axis Hydrogen bond