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3.1 - Stoichiometry

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Match. Gravity. Created by.
rwilliams1. Terms in this set (7)
actual yield. the amount of
product that forms when a
reaction is carried out in the
laboratory. excess reagent.

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Answer Key Stoichiometry The

study of quantitative relationships

between the amounts of

reactants used and amounts of

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products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass.

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Stoichiometry 1. $1 \text{ ZnI}_2 \rightarrow 1 \text{ Zn} +$
 1 I_2 How many grams of iodine
will you produce if you begin your

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rxn with 56.7g of zinc iodide?

56.7 g 1 mole 1 mole 253.8 g

319.19 g 1 mole 1 mole = 45.08

grams of Iodine 2. 1 $\text{Pb}_2(\text{CO}_3)_3 +$

6 K \rightarrow 2 Pb + 3 K_2CO_3

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1. $1 \text{ ZnI}_2 \rightarrow 1 \text{ Zn} + 1 \text{ I}_2$ How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide?

56.7 g 1 mole 1 mole 253.8 g
319.19 g 1 mole 1 mole = 45.08
grams of Iodine

2. $1 \text{ Pb}_2(\text{CO}_3)_3 + 6 \text{ K} \rightarrow 2 \text{ Pb} + 3 \text{ K}_2\text{CO}_3$

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rxn with 56.7g of

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